

Nanocomposites thin films as gas sensor application

J. M. Patil ^{*a}, S. B. Patil ^b R. H. Bari^c, and A. N. Sonar ^a

^{a*}Department of Chemistry, Shri. V. S. Naik, A.C.S. College, Raver, 425508, Maharashtra, India.

^{b*}Department of Physics, A.R.B.Garud ACS College,Shendurni.

^cNanomaterials Research Laboratory, Department of Physics, G. D. M. Arts, K. R. N.

Com. And M. D. Science College, Jamner, 424 206, Maharashtra, India.

Abstract

In present work nanocomposites thin films were prepared by spray pyrolysis technique. As prepared thin films were studied using XRD and FE-SEM and. The film sprayed for composition of $WO_3-V_2O_5$ (Sample =S2) was observed to be most sensitive ($S = 1130$) to SO_2 at $350\text{ }^\circ\text{C}$. The sensor shows quick response (4 s) and fast recovery (8 s) time. The results are discussed and interpreted.

Keywords: Spray Pyrolysis, $WO_3-V_2O_5$ nanocomposites, SO_2 gas sensing, quick response, fast recovery

1. Introduction

Metal oxide semiconductors are used extensively as a sensing element of different gases and vapors. A depletion region always formed at the surface of metal oxide semiconductor due to adsorption of air oxygen molecules. Then the reaction with the target gas molecule causes reduction of depletion region which results change in conductivity of metal oxide semiconductor. The conductivity may increase or decrease depending on type of semiconductor and type of target gas [1]. The metal oxide-sensing layer (WO_3 or V_2O_5) has been fabricated in different physical forms such as thin film, thick films, and bulk pellets. However, the thin film form is expected to be most effective, because sensing is basically a surface phenomenon of film. Thus, a very few work has been reported for the combination of WO_3/V_2O_5 oxide composite.

In present work efforts was done in the area of SO_2 detection using metal oxide thin films . However, not much attention has been given to the fabrication of nanocomposites structure for detection of SO_2 gas. There has been intensive research on improving the gas sensitivity and selectivity by controlling the particle size, nanostructures, sensing temperature, surface and structure [2].

2. Experimental details

2.1 Preparation of WO_3 - V_2O_5 nanocomposites thin films

The starting material used for the preparation of WO_3 - V_2O_5 nanocomposites thin films were tungsten hexachloride (WCl_6 , Purified Merck) and vanadium (III) chloride (VCl_3 , Purified Aldrich). Tungsten hexachloride and vanadium (III) chloride were mixed at various volume ratio such as 30:70, 50:50 and 70:30 as indicated in Table 1.

Table 1: Amounts of spraying solutions and reactant

Sample No.	WCl_6 (cm^3)	VCl_3 (cm^3)	Reactants
S1	30	70	WO_3 : V_2O_5
S2	50	50	WO_3 : V_2O_5
S3	70	30	WO_3 : V_2O_5

The optimized deposition parameters like substrate temperature (350 °C), spray time (10 mn.), rate of spraying solution (8 ml/min.), nozzle to substrate distance (30 cm), quantity of the solution sprayed (30 ml), pressure of carrier gas, and to and fro movement of the nozzle were kept constant. The temperature of the substrate is maintained at a constant value by using a temperature controlled hot plate. The film formation depends upon the droplet landing, reaction and solvent evaporation, which relates to the droplet size. When the droplet approaches the substrate just before the solvent is completely removed, that is the ideal condition for the preparation of the pure WO_3 , V_2O_5 , and WO_3 - V_2O_5 nanocomposites thin film. The as prepared WO_3 , V_2O_5 , and WO_3 - V_2O_5 nanocomposites thin films samples were annealed at 500 °C for 1 h.

3. Results and discussion

3.1. Structural analysis using X-ray diffractogram

Fig. 1 shows the X-ray diffractogram of thin film sample. The observed peak predominates indicating a preferential growth. This means that the grains have c -axis perpendicular to the substrate surface. The calculated average crystallite size was found to be 18 nm.

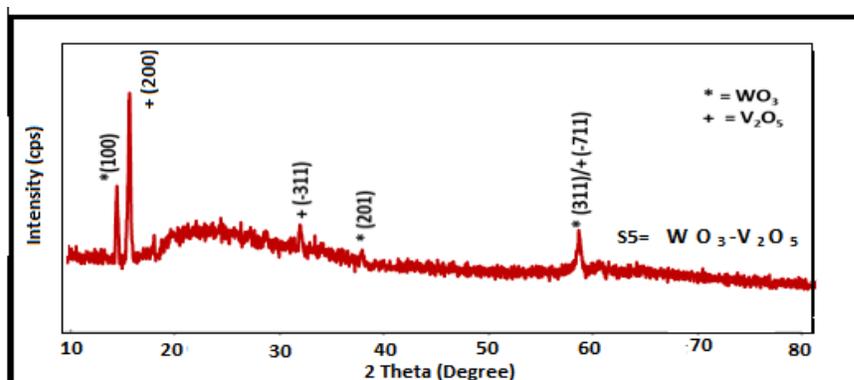


Fig. 1. X-ray diffractogram of most sensitive nanocomposites thin film (sample S2)

3.2. Field emission scanning electron microscope

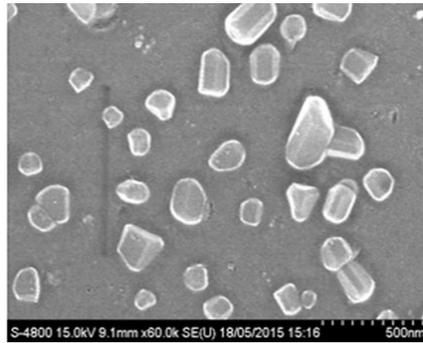


Fig.2. FE-SEM images of most sensitive nanocomposites thin film (sample S2)

FE-SEM images of pure WO_3 , V_2O_5 , and $\text{WO}_3\text{-V}_2\text{O}_5$ nanocomposites were represented in Fig.2. Grain size observed to be in the range of 21 - 44 nm.

4. Gas sensing performance of the sensors

4.1. Gas response

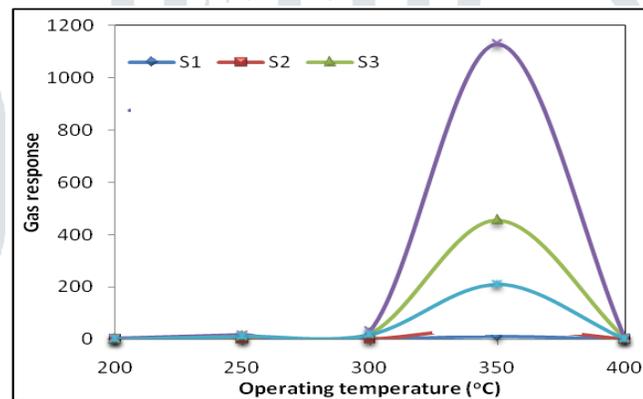


Fig. 3. Gas response of nanocomposites thin films with operating temperature

Fig. 3 represents the response characteristics of the $\text{WO}_3\text{-V}_2\text{O}_5$ nanocomposite thin films as a function of operating temperature. Among all the films, the sample (S2) film shows the maximum response (1130) at 350 °C to 500 ppm of SO_2 .

4.2. Response and recovery of the sensor with concentration of gas in ppm

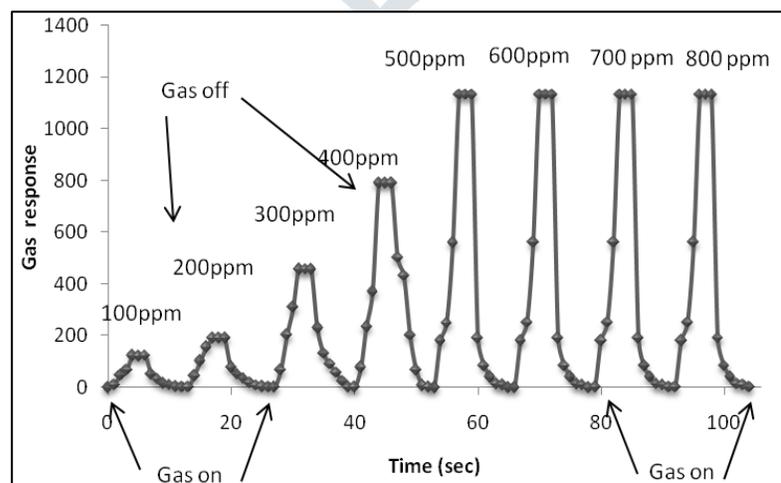


Fig. 4 Response and recovery of the sensor for most sensitive sample (S2).

The response is quick (4 s) and recovery is fast (8 s). The high oxidizing ability of adsorbed oxygen species on the surface nanoparticles and high volatility of desorbed by-products explain the quick response to H₂S and fast recovery [3-5].

5. Conclusion

Nanocomposites thin films were prepared by simple spray pyrolysis technique. The WO₃-V₂O₅ thin film of (Sample S2) was most sensitive to SO₂ gas and exhibits the response of S = 1130 to the gas concentration as 500 ppm at the temperature of 350 °C. The WO₃-V₂O₅ nanocomposites thin films exhibit rapid response–recovery which is one of the main features of this sensor.

Acknowledgements

The authors are thankful to the University Grants Commission, New Delhi for providing financial support. Thanks to Principal, Shri. V. S. Naik A.C.S. College, Raver, for providing necessary infrastructure laboratory facilities for this work. I have also appreciate Prof. Dr. R. H. Bari, Head of Department, Department of Physics, G. D. M. Arts, K. R. N. Com. And M. D. Science College, Jamner for their help. We thanks NMU, Jalgaon for the motivation and encouragement for giving me the opportunity to do this research work as successful one.

Reference

- [1] P. J. Shaver, Appl. Phys. Lett., vol. 11, pp. 255-256, 1967.
- [2] U. Schlecht, I. Besnard, A. Yasuda, T. Vossmeier, M. Burghard, in: H. Kuzmany, J. Fink, M. Mehring, S. Roth (Eds.), American Institute of Physics, Melville, NY, 2003, p. 491.
- [3] J. Muster, G.T. Kim, V. Krstic, J.G. Park, Y.W. Park, S. Roth, M. Burghard, , Adv. Mater. 12 (2000) 420–424.
- [4] R. H. Bari, S. B. Patil, A.R. Bari, G. E. Patil, J. Aambekar, Sensors & Transducers Journal 140, 124 (2012)
- [5] L.A. Patil, A.R. Bari, M.D. Shinde, Vinita Deo, Sensors and Actuators B: Chemical, **149**, 79 (2010)